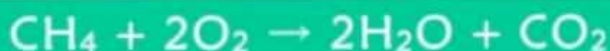


## COMBUSTION



**COMBUSTION** or burning is a high temperature exothermic redox chemical reaction between a fuel and an oxidant usually atmospheric oxygen.



## ADDITION

**ADDITION** reaction is a reaction in which one molecule combines with another to form a large molecule with no other product.



## DECOMPOSITION

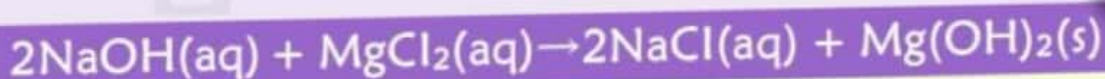


**DECOMPOSITION** reaction involves the breakdown of chemical compound into its elements or simpler compounds. These reactions often involve an energy source such as heat, light, or electricity that breaks apart the bonds of compounds.



## PRECIPITATION

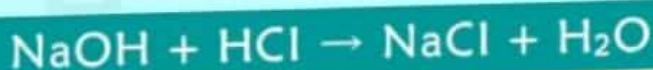
**PRECIPITATION** reaction is the one in which aqueous compounds react to form an insoluble solid, called a precipitate. Whether or not a reaction will form a precipitate is dictated by solubility rules for ionic compounds.



## NEUTRALISATION



**NEUTRALISATION** reaction is the one in which an acid and a base react to form a salt and water. Neutralisation reactions do not necessarily result in a pH of 7, resultant pH is dependent on the strengths of the acid and base.





# CHEMICAL REACTIONS

## CONDENSATION



**CONDENSATION** reaction is the one in which two molecules combine to form a larger molecule with a small molecule. The small molecule lost is usually water, but not always. It can be considered to be the opposite of hydrolysis.



## HYDROLYSIS

**HYDROLYSIS** reaction involves the breaking of chemical bonds by the addition of water to a substance. In some cases, this addition can cause both the substance and the water molecule to split into two parts.



## DISPLACEMENT

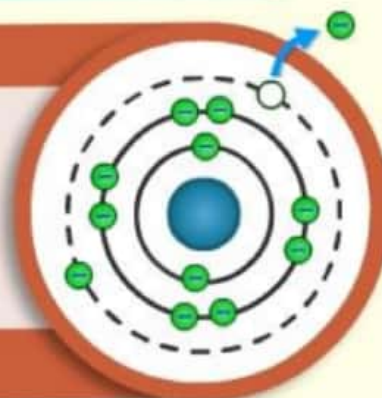
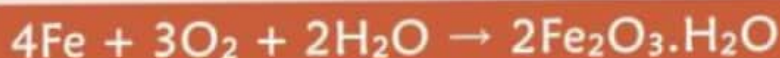


**DISPLACEMENT** reaction is chemical reaction in which a more reactive element displaces a less reactive element from its compound. Both metals and non-metals take part and more reactive displaces the less reactive.



## OXIDATION

**OXIDATION** reactions are sometimes defined as the reactions in which an element forms bonds with oxygen atoms. Generally, oxidation can be defined as a reaction in which atoms of an element **lose electrons**.



## REDUCTION

**REDUCTION** reactions are sometimes defined as reactions in which other molecules lose oxygen atoms. Generally, reduction can be defined as a reaction in which atoms of an element **gain electrons**.

